
A comparative study of metachromasy induced in Azure B by anionic polyelectrolytes

Nandini Ratnakar

Department of Chemistry, MITE, Moodabidri-574 226 (DK), Karnataka, India

Email address:

nandinifalnr@yahoo.com

To cite this article:

Nandini Ratnakar. A Comparative Study of Metachromasy Induced in Azure B by Anionic Polyelectrolytes. *Science Journal of Chemistry*. Special Issue: Polyelectrolytes & Dyes. Vol. 2, No. 6-2, 2014, pp. 8-13. doi: 10.11648/j.sjc.s.2014020602.12

Abstract: The interaction of two cationic dyes, namely, Azure B (AB) with an anionic polyelectrolyte, namely, sodium heparinate (NaHep) has been investigated by Spectrophotometric method. The polymer induced metachromasy in the dyes resulting in the shift of the absorption maxima of the dyes towards shorter wavelengths. The stability of the complexes formed between Azure B and sodium alginate was found to be greater than that formed between Azure B and sodium carboxymethylcellulose. This fact was further confirmed by reversal studies using alcohols, urea surfactants and electrolytes. The interaction parameters revealed that binding between Azure B and anionic polyelectrolytes was mainly due to electrostatic interaction.

Keywords: Cationic Dye, Metachromasy, Anionic polyelectrolytes, Electrostatic Interaction, Aggregation

1. Introduction

Metachromasy is a well-known phenomenon in the case of dye-polymer interactions and is generally found in the case of aggregation of cationic dyes on anionic polymers [1-4]. Metachromasy is related to the interaction of cationic dyes with polyanions where a single individual compound is formed by the interaction of the dye cation and the chromotrope polyanionic polymer. Several physicochemical parameters such as molecular weight of each repeating unit, stoichiometry of the dye-polymer complex, binding constant, and other related thermodynamic parameters like free energy, enthalpy, and entropy changes can be evaluated using polymer-dye interactions. Biological activity of macromolecules depends on its tertiary conformation. Conformation of the polyanions controls the induction of metachromasy of aqueous dye solution. Although there are several reports on metachromasy of various classes of acidic polysaccharides and different synthetic polyanions [5, 6]. Cyanine dyes, which are cationic in nature, have widely been used to probe biological systems such as helical structure of DNA, tertiary conformation of bacterial polysaccharides and other polymers. As these dye having high light absorptivity, they can be used as optical probes in studying membranes, micelles and other host systems [7-10]. Studies on polymer-surfactant interaction in aqueous solution have been

attracting widespread attention due to multiple practical uses in biology [11-14]. Such studies are also assumed to be important as the mixed systems/aggregates can give rise to advanced functions that are unobtainable from single component [15]. Several physicochemical properties of macromolecule - surfactants are quite relevant in this context. Formulation procedures based on suitable mixture may have been appealing applications [16-18]. It has been noted that oppositely charged surfactant binds to polymer surfaces through both electrostatic and hydrophobic interaction [19]. Different techniques for their isolation and stability determination of metachromatic complexes have been reported in literature [20]. The phenomena of reversal of metachromasy by addition of urea, alcohol neutral electrolytes [21] and also by increasing the temperature of the system may be used to determine the stability of the metachromatic compounds. The nature of dye-polymer interaction in the metachromatic complex formation and also the conditions for the interaction between the cationic dye and the anionic site of the macromolecules has been studied by determining the thermodynamic parameters of the interaction [22]. The interaction between polyelectrolytes and oppositely charged surfactants [23,24] has been investigated due to its importance in both fundamental and applied fields. [25]. Comprehensive studies of a variety of cationic surfactants with anionic polyelectrolytes, have been reported [26-32]. Cationic dyes that interacted with Heparin mostly fell

in the Thiazine group such as MethyleneBlue [33-35]. Azure A [36, 37], Azure B [38,39]. Hence the objective of the present study is to compare the extent of metachromasy induced by sodium alginate and sodium carboxymethyl cellulose in the cationic dyes Azure B and to evaluate the thermodynamic parameters of interaction and to study the extent of reversal by using alcohols, urea, surfactants and electrolytes which is an indirect evidence for the stability of the metachromatic complex formed. The effect of dye structure on metachromasy has also been discussed.

2. Materials & Methods

2.1. Apparatus

The spectral measurements were carried out using Shimadzu UV-2500 Spectrophotometer using a 1cm quartz cuvette.

2.2. Reagents

Azure B (AB) was obtained from (Hi Media, Germany) and used as received.

Sodium carboxymethyl cellulose (NaCMC) & sodium heparinate (NaHep) (Loba Chemie, India) were used without further purification; Methanol, (MeOH) ethanol (EtOH) and 2-Propanol (PrOH) (Merck, India) were distilled before use. Urea, sodium chloride and potassium chloride were obtained from (Merck, India). sodium laurylsulphate and sodium dodecylbenzenesulphonate (Lobachemie, India) were obtained from (Loba Chemie, India).

3. Methods

3.1. Determination of Stoichiometry of Polymer-Dye Complex

Increasing amounts of polymer solution (0.0- to 5×10^{-3} M) were added to a fixed volume of dye solution (0.5ml, 1×10^{-5} M) in case of Azure B in different sets of experiments and the total volume was made up to 10ml by adding distilled water in each case. The absorbances were measured at 547nm in case of AB-NaHep and at 562 nm in case of AB-NaCMC.

3.2. Study of Reversal of Metachromasy Using Alcohols and Urea

For measurements of the reversal of metachromasy, solutions containing polymer and dye in the ratio 2.0:1.0 were made containing different amount of alcohol. The total volume was maintained at 10ml in each case. The absorbances were measured at 645 nm and 545 nm in case of AB-NaHep and at 645 nm and 562 nm in case of AB-NaCMC. Similarly, polymer-dye solutions containing different amounts of urea (1-8M) were made and the absorbances were measured at 645 nm and at 547 nm in case of AB-NaHep complex and at 562 nm in case of AB-NaCMC complex.

3.3. Study of Reversal of Metachromasy Using Surfactants and Electrolytes

For measurements of the reversal of metachromasy, solutions containing polymer and dye in the ratio 2.0:1.0 were made containing different amount of surfactants or (1×10^{-7} M – 1×10^{-2} M) were made in case of AB-NaHep and (1×10^{-6} M-0.1M) in case of AB-NaCMC. The total volume was maintained at 10ml in each case. The absorbances were measured at 645 nm and 547 nm in case of AB-NaHep and at 645 nm and 562 nm in case of AB-NaCMC. Similarly, polymer-dye solutions containing different amounts of electrolytes (1×10^{-6} M-0.1M) in case of AB-NaAlg and (1×10^{-7} M- 1×10^{-2} M) in case AB-NaCMC were made and the absorbances were measured at 645 nm and at 547 nm in case of AB-NaHep complex and at 645 nm and at 562 nm in case of AB-NaCMC complex. A plot of absorbance at 645 nm Vs molar concentration of surfactants or electrolytes was made to determine the concentration of surfactants and electrolytes needed for complete reversal of metachromasy.

3.4. Determination of Thermodynamic Parameters

The thermodynamic parameters were determined by measuring the absorbances of the pure dye solution at the respective monomeric band and metachromatic band in the temperature range (360C-540C). The above experiments were repeated in presence of the polymers at various polymer-dye ratios (2, 5, 8 and 10).

4. Results & Discussion

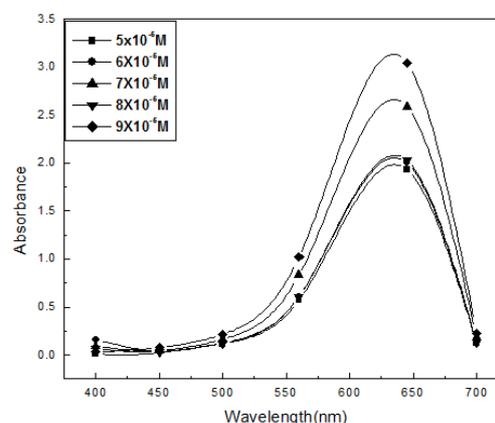


Figure 1. Absorption spectrum of AB at various concentrations

The absorption spectra of Azure B shows an absorption maximum at 645 nm in case of Azure B indicating the presence of a monomeric dye species in the concentration range studied. On adding increasing amounts of polymer solution the absorption maxima shifts to 547 nm in case of AB-NaHep and at 562 nm in case of AB-NaCMC complex. The blue shifted band is attributed to the stacking of the dye molecules on the polymer backbone and this reflects high degree of co-operativity in binding [40]. Appearance of multiple banded spectra proposed that the polymer might have a random coil structure in solution. Whereas at higher

concentration of the polymer almost a single banded spectrum was observed due to possible change from random coil to helical form [41]. The absorption spectra at various P/D ratios are shown in (Fig.1 and Fig.2) respectively.

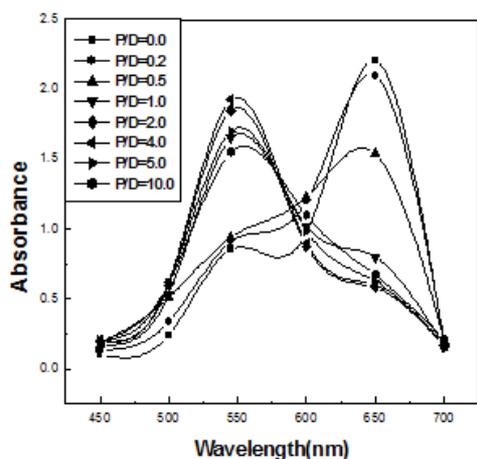


Figure 2. Absorption spectrum of AB-NaHep at various P/D ratios

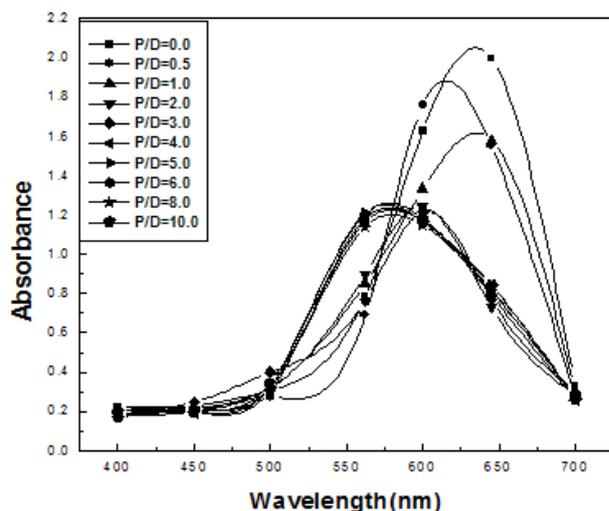


Figure 3. Absorption spectra of AB-NaCMC system at various P/D ratio

4.1. Determination of Stoichiometry

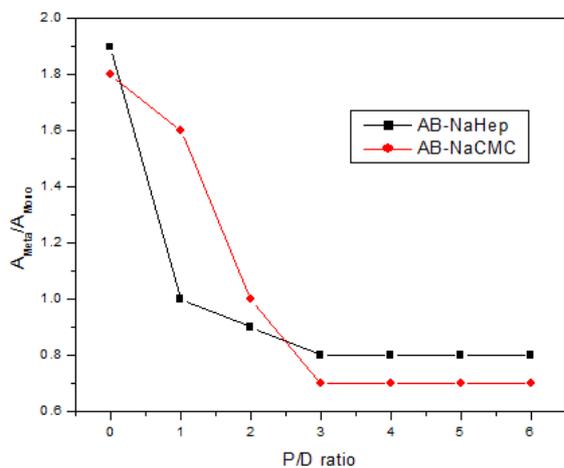


Figure 4. Stoichiometry of AB-NaHep & AB-NaCMC complex

To determine the stoichiometry of the polymer-dye complex, a plot of A_{547}/A_{645} versus the polymer/dye ratio was made for AB-NaHep system. A similar procedure was repeated with AB-NaCMC complex also. The stoichiometry of AB-NaCMC complex was found 2:1 which indicates that the binding is at alternate anionic sites. This indicates that every potential anionic site of the polyanion was associated with the dye cation and aggregation of such dye molecules was expected to lead to the formation of a card pack stacking of the individual monomers on the surface of the polyanion so that the allowed transition produces a blue-shifted metachromasy [42]. The results were in good agreement with the reported values for interaction of similar dyes with polyanions [43]. While in case of AB-NaAl complex the stoichiometry is 1:1 and the binding is at adjacent anionic sites. This indicates that there is lesser overcrowding and more aggregation of the bound dyes on the polymer chain in the latter case than in the former case. Similar results were reported in case of binding of Pinacyanol chloride on poly(methacrylic acid) & poly(styrenesulfonate) systems [44]. The stoichiometry results are obtained by plotting A_{645}/A_{547} or A_{645}/A_{562} Vs P/D ratio in each case. The results are shown in (Fig. 4).

4.2. Reversal of Metachromasy Using Alcohols and Urea

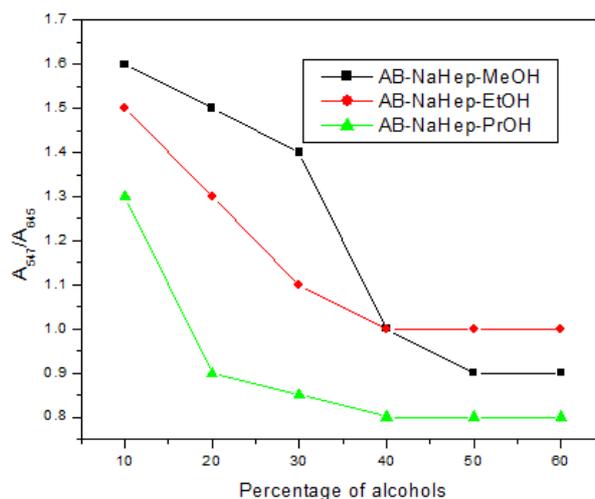


Figure 5. Reversal of metachromasy on addition of alcohols (AB-NaHep)

The metachromatic effect is presumably due to the association of the dye molecules on binding with the polyanion which may involve both electrostatic and hydrophobic interactions. The destruction of metachromatic effect may occur on addition of low molecular weight electrolytes, alcohols or urea. The destruction of metachromasy by alcohol and urea is attributed to the involvement of hydrophobic bonding has already been established [45-49]. The efficiency of alcohols in disrupting metachromasy was found to be in the order methanol < ethanol < 2-propanol, indicating that reversal becomes quicker with increasing hydrophobic character of the alcohols. The above facts are further confirmed in the present system. On addition of increasing amount of alcohol

to the polymer/dye system at P/D=2.0, the original monomeric band of dye species is gradually restored. The efficiency of the alcohols, namely methanol, ethanol and 2-propanol, on destruction of metachromasy were studied. In case of AB-NaHep system 60 % methanol, 50 % ethanol, 40 % 2-Propanol were sufficient to reverse metachromasy 40 % methanol, 30 % ethanol, 20 % 2-Propanol were required to reverse metachromasy in AB- NaCMC system. From the plot of A_{547}/A_{645} or A_{562}/A_{645} (Fig.5 and Fig6) against the percentage of alcohols or molar concentration of urea, the percentage of alcohols or molar concentration of urea, needed for complete reversal has been determined. The concentration of urea to reverse metachromasy is found to be as high as 4M in AB-NaHep system and 3M in case of AB-NaCMC system (Fig7) This indicates that the metachromatic complex formed between AB-Na Alg is more stable than that between AB-NaCMC complexes. Similar reports are available in literature for reversal of metachromasy in anionic polyelectrolyte/cationic systems by addition of alcohols or urea [50-51].

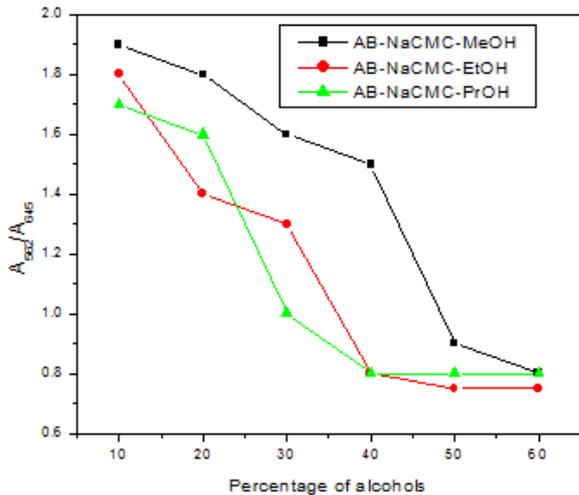


Figure 6. Reversal of metachromasy on addition of alcohols (AB-NaCMC)

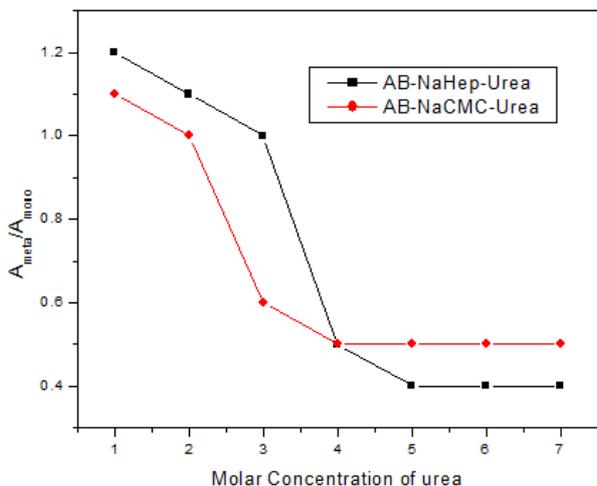


Figure 7. Reversal of metachromasy on addition of Urea in AB-NaHep & AB-NaCMC systems

4.3. Reversal of Metachromasy Using Surfactants

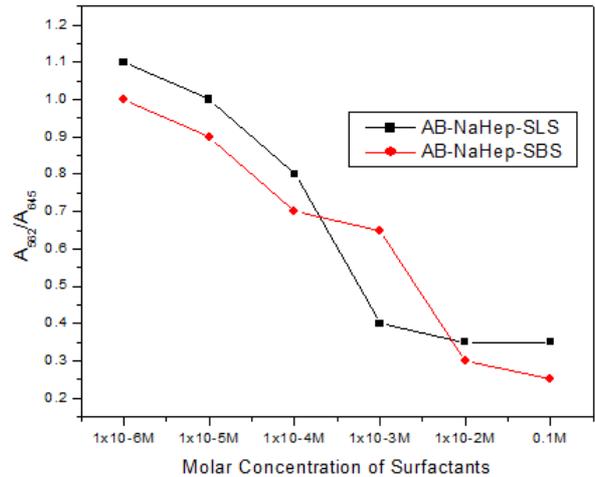


Figure 8. Reversal of metachromasy on addition of surfactants-ABNaHep

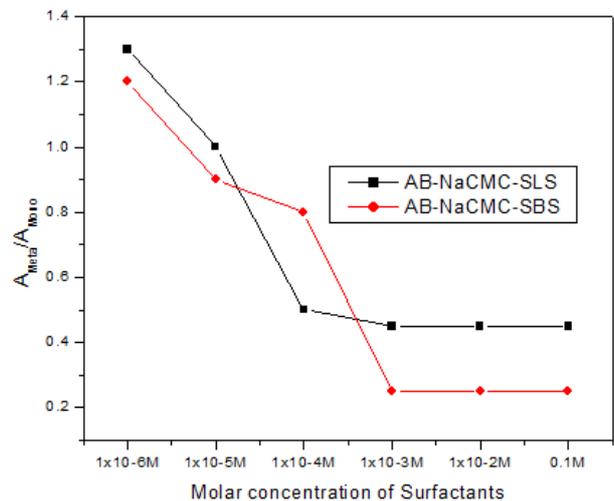


Figure 9. Reversal of metachromasy on addition of surfactants (AB-NaCMC)

The strength and nature of interaction between water soluble polyelectrolytes and oppositely charged surfactants depend on the characteristic features of both the polyelectrolytes and the surfactants. The charge density, flexibility of the polyelectrolyte and the hydrophobicity of the non-polar part and the bulkiness of the polar part also play a vital role in the case of polysaccharide-surfactant interaction [52]. On adding increasing amounts of sodium laurylsulphate and sodium dodecylbenzenesulphonate to AB-NaHep complex the molar concentrations of Sodium laurylsulphate and Sodium dodecylbenzenesulphonate needed to cause reversal was found to be $1 \times 10^{-3} M$ and $1 \times 10^{-2} M$ in case of AB-NaHep and $1 \times 10^{-3} M$ and $1 \times 10^{-3} M$ in case of AB-NaCMC system. These results agree with those reported earlier in literature [53,54]. Thus the addition of surfactants causes the production of micelles. Thus the surfactant molecules interacted with the polymer by replacing the cationic dye. The release of dye molecules from the dye-polymer complex in presence of cationic surfactants revealed that surfactants interacted electrostatically [55,56] with the

anionic site of the polymer and thus the dye becomes free. The ease of reversal of metachromasy can be correlated with its chain length thus the binding between oppositely charged polymer surfactant is primarily by electrostatic forces which is reinforced by hydrophobic forces. From the plot of absorbance at $A_{\text{meta}}/A_{\text{mono}}$ in case of Azure B against molar concentration of surfactants, the molar concentration of surfactants needed for complete reversal of metachromasy has been determined. The results are shown in (Fig. 8 and Fig.9)

The interaction constant K_C for the complex formation between AB and NaHep and AB-NaCMC was determined by absorbance measurements at the metachromatic band at four different temperatures taking different sets of solutions containing varying amounts of polymer (C_S) in a fixed volume of the dye solution (C_D). The value of K_C was obtained from the slope and intercept of the plot of $C_D C_S / (A - A_0)$ against C_S shown in Fig.10 & Fig.11. In each case the thermodynamic parameters of interaction, namely ΔH , ΔG and ΔS were also calculated. The results are given in Table 1. (C_S in a fixed volume of the dye solution (C_D)). The value of K_C was obtained from the slope

Table 1. Effect of temperature on AB-NaHep & AB-NaCMC system

System	Temp (K)	K_C ($\text{dm}^3 \text{mol}^{-1}$)	ΔG ($\text{kJ} \cdot \text{mol}^{-1}$)	ΔH ($\text{kJ} \cdot \text{mol}^{-1}$)	ΔS ($\text{J} \cdot \text{mol}^{-1} \text{K}^{-1}$)
AB-NaHep	309	10783	-23.8		
	315	8095	-23.7	-18.0	-28.3
	321	7628	-23.5		
	327	6324	-23.4		
AB-NaCMC	309	9856	-23.8		
	315	8176	-24.0	-26.3	-33.3
	321	7654	-24.2		
	327	5432	-24.3		

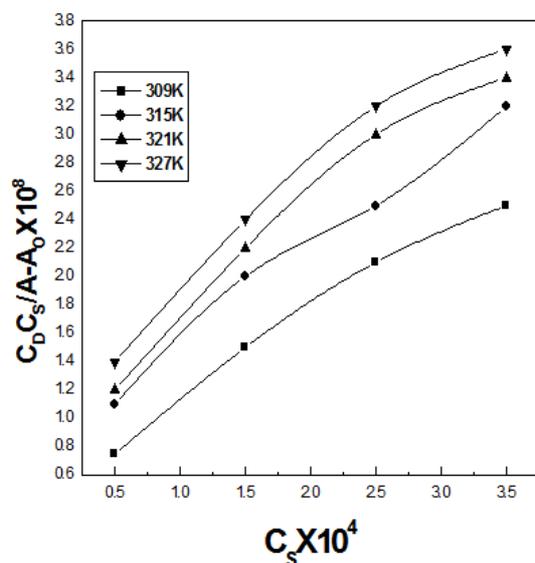


Figure 10. Effect of temperature on AB-NaHep system

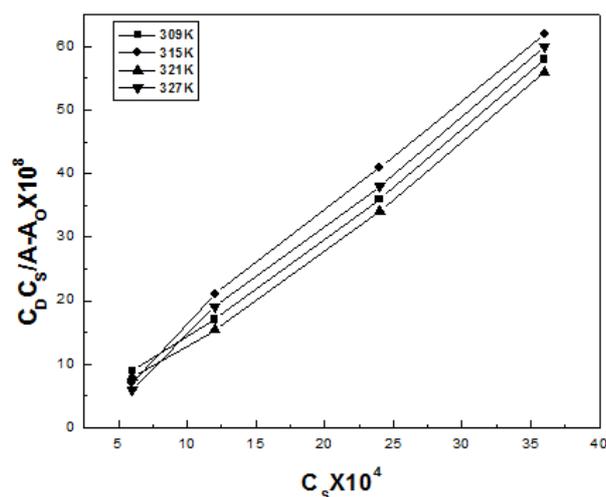


Figure 11. Effect of temperature on AB-NaCMC system

5. Conclusion

The polymers sodium heparinate and sodium carboxymethyl cellulose induced metachromasy in the dye Azure B. The monomeric band occurs at 645 nm while the metachromatic band occurred at 547 nm in the case of AB-NaHep (98nm) and at 562nm in the case of AB-NaCMC.(83nm). These results are further confirmed by the reversal studies using alcohols, urea, electrolytes and surfactants. The thermodynamic parameters further confirm the above fact. It is thus evident from the above studies that both electrostatic and non-ionic forces contribute towards the binding process. Based on the results it can be concluded that the Azure B is more effective in inducing metachromasy in sodium heparinate than in sodium carboxymethyl cellulose. Cooperativity in binding is observed to occur due to neighbour interactions among the bound dye molecules at lower P/D ratios leading to stacking. The stacking tendency is enhanced by the easy availability and close proximity of the charged sites.

References

- [1] A Mitra, R. K Nath., SBiswas., A. K., Chakraborty, A. K., Panda Mitra A., J. Photochem. Photobiol. A: Chem, vol.111, 1997, pp 157.
- [2] A.Mitra, A.K.Chakraborty, J.Photochem.Photobiol.A:Chem., vol.178,2006, pp198.
- [3] R.W Horbin., Biochemie.Biochem., vol.77, 2002, pp 3.
- [4] Bergeron J. A., M. J Singer., Biophys. Biochem. Cytol. vol.4,1958, pp.433.
- [5] M. K. Pal, B. K. Ghosh, Makromol. Chem. vol. 181, 1980, pp1459.
- [6] Norden B., Kubista M., in : B.Samori, E.W.Thulstrup (Eds.), Polarized Spectroscopy of ordered systems, vol.242,Kulwer Academic Publisher, Dordrecht, Holland, 1988.
- [7] Sabate R., EsterlichJ., J.Phys.ChemB, vol.107,2003, pp.4137.

- [8] von Berlepsch H., Kirstein S., Bottcher, C, Langmuir, vol.18,2002,pp.769.
- [9] Sabate R., Gallardo M., De la Maza A., J.Esterlich, Langmuir,vol.17,2001, pp.6433.
- [10] Berret J.F., Cristobal G.,Hervel P., Oberdisse J.,Grillo I., vol.E9, Eur.Phys J 2002, pp.301.
- [11] Meszaos R., Varga, I.,Gilanyi T.,J.Phys.Chem.B vol.117, 2005, pp.13538.
- [12] Monteux C . ,WilliamsC.E., Meunier J., Anthony O., BergeronV., Langmuir Vol.20,2004,pp. 57.
- [13] Honda C., Kamizono H.,Matsumoto K., Endo K., J.Colloid Interface Sci vol.278.2004, , pp.310.
- [14] MoulikS.P., GuptaS.,DasA.R., Makromol.Chem,vol.181, 1980,pp. 1459.
- [15] Lee J ., Moroi Y.,J.Colloid Interface Sci. 2vol.273,2004, pp.645.
- [16] Mesa C.L., J.Colloid Interface Sci . vol.286,2005, pp.148.
- [17] Konradi R.,Ruhe J., Macromolecules vol.38,2005, pp.6140.
- [18] VillettiM., BorsaliA., Crespo R.,SoldiV.,Fukada K., Macromol.Chem.Phys. vol.205,2004, pp.907.
- [19] Fundin J.,Hansson P.,Brown W., Lidegran I.,Macromolecules vo.30, 1997, pp.1118.
- [20] Mitra A., Nath A. R., and ChakrabortyA.K.,Colloid Polym Sci. , vol.271,1993, pp.1042.
- [21] Panda A.K., Chakraborty A.K., J.Colloid Interface Sci. vol.203,1998, pp.260.
- [22] Chakraborty, A.K.,Nath, R.K, Spectrochim Acta vol.45A,1989, pp.981.
- [23] Jain N. ,Trabelsi S., Guillot S.,Meloughlin D. ,Langevin D., Letteiller P., Turmine, M., Langmuir, vol..20,2004, pp.8496.
- [24] Bakshi M.S., Varga I., GilanyiT.,J.Phys.Chem.B,vol. 109 ,2005, pp.13538.
- [25] Balomenou L.,and Bokias G., Langmuir, vol. 21,2005, pp.9038.
- [26] Bakshi M.S., and Sachar S.,Colloid Polym.Sci. vol.283,2005,pp.671.
- [27] Romani A.P., Gehlen M.H., Itri R., Langmuir, vol.21,2005, pp.127.
- [28] Wang C., Tam K.C., Langmuir, vol.18,2002, pp. 6484.
- [29] Sjogren H .,Ericsson C.A., Evenas J.,Ulvenlund S .,Biophys J., vol.89,2005, pp.4219 .
- [30] Zhu D.M., EvansR.K., Langmuir, 2006, 22 ,pp.3735.
- [31] Chatterjee A.,Moulik S.P., Majhi P.R., and Sanyal S.K., Biophys Chem.vol.98,,2002,98, pp.313.
- [32] Mata J., Patel J., Jain N.,Ghosh G.,Bahadu, P., J.Colloid Interface Sci. 2006, 297 , pp.797.
- [33] Jiao Q.C., Liu Q .,Sun C.,and He H ., Talanta, vol.1095,1999, , pp.48.
- [34] Liu Q.,and Jiao Q.C .,Spectrosc.Lett., vol.31,1998, pp.913.
- [35] Jiao Q.C. ,and Liu Q.,Spectrochim. Acta, 1vol.55A,999, pp.1667.
- [36] Jiao Q.C and Liu Q., Anal.Lett. vol..31,1998, pp.1311.
- [37] Jiao Q.C. ,and Liu Q, Spectrosc.Lett.vol.31, 1998, pp.1353.
- [38] Hugglin D., Seiffert, A.,Zimmerman, W,Histochemistry, vol.86,1986, pp.71.
- [39] Mitra A ., Nath R.K .,and. Chakraborty A.K., Colloid Polym. Sci., vol.271,1993, pp.1042.
- [40] Basu. S. , Gupta A.K. ,and Rohatgi-Mukherjee K. K, J.Indian Chem. Soc.1vol.59,982, pp.578 .
- [41] Pal M.K., SchubertM.J. Histochem Cytoche., vol.9,1961,pp. 673.
- [42] Pal M.K.,and Ghosh B.K., Macromol.Chem, vol.181,1980,pp. 1459.
- [43] Pal M.K and Ghosh B.K., Macromol.Chem. vol.180,1979, pp.959.
- [44] Pal M.K and Schubert M .,J.Phys.Chem., vol.17,1963,pp.182.
- [45] Frank. H.S., and.Evans M.W .,J.Chem.Phys. vol.17,1945, pp.507 .
- [46] Kauzmann W.,Advan.Protein.Chem.vol.14, 1959, pp1.
- [47] Bruning and Holtzer A., J.Am.Chem Soc.,vol.83,1961, pp.4865.
- [48] WhitneyP.L., and Tanford C., J.Biol.Chem vol.237,1962, pp.1735 .
- [49] Mukerjee P.,and Ray A. ,J.Phys.Chem. vol.67,1963, pp.190.
- [50] Frank H.S. and Quist A.S., J.Chem.Phys.,1vol.34,961, pp.604.
- [51] Rabinowitch E., Epstein., L.F J. Am. Chem. Soc., vol.63, 1941,pp.69.
- [52] Romani A.P., Gehlen M.H. and. Itri R., Langmuir, vol.21,2005, pp.127.
- [53] Levine A., Schubert M., J.Am.Chem.Soc., vol.74,1958, pp.5702.
- [54] Dasgupta S., Nath R.K. ,Biswas S., Hossain J., Mitra A.,and Panda A.K. ,Colloids Surf A: Physicochem Eng.Aspects, vol.302,2007, pp.17.
- [55] RoseNJ; DragoRS; Journal of American Chemical Society, 1vol.81,959, pp.6138.
- [56] Nandini R, Vishalakshi B,e-journal of chemistry, vol.8,2011, pp.S253.
- [57] Nandini R, Vishalakshi B,spectrochimica acta Part A,, vol.9,2009, pp.1025