



Crystal Structure Imperfection of LiFePO_4 Synthesized Through Solid-state Reaction: An XRD Overview

Ibrahim Purawiardi^{1, *}, Christin Rina Ratri²

¹Research Centre for Metallurgy and Materials, Indonesian Institute of Sciences (LIPI), South Tangerang, Indonesia

²Research Centre for Physics, Indonesian Institute of Sciences (LIPI), South Tangerang, Indonesia

Email address:

rade026@lipi.go.id (I. Purawiardi), ibrahimpurawiardi@gmail.com (I. Purawiardi), christin.rina.ratri@lipi.go.id (C. R. Ratri)

*Corresponding author

To cite this article:

Ibrahim Purawiardi, Christin Rina Ratri. Crystal Structure Imperfection of LiFePO_4 Synthesized Through Solid-state Reaction: An XRD Overview. *International Journal of Mineral Processing and Extractive Metallurgy*. Vol. 1, No. 5, 2020, pp. 1-6.

doi: 10.11648/j.ijmpem.20200501.11

Received: March 5, 2020; **Accepted:** March 19, 2020; **Published:** March 31, 2020

Abstract: Solid-state reaction is one of some methods to synthesize LiFePO_4 powder. However, the post-synthesis crystal structure was found to be imperfect, probably caused by the imperfection of the phospho-olivine structure. This study aimed to investigate the cause of its imperfection. A LiFePO_4 powder synthesized via solid-state reaction path was used as a case study sample for this study. XRD characterization was done to investigate it. Orthorhombic crystal structure was found to be a perfect fit for this sample using precise lattice parameter analysis, as shown by the linear regression equation result. Further analysis was performed using Rietveld refinement method to pinpoint the actual coordinates of Li, Fe, P, and O atoms. The result shows that solid-state reaction can produce an order of orthorhombic crystal structure which constructed by ordered Li atoms arrangement. On the other hands, there is a disordered phospho-olivine structure due to the imperfection of the occupation of Fe, P, and O atoms. These disorders were found through analysis of anomalous peaks on the Rietveld refinement result when compared with PDF database. Loss of (200) plane was caused by imperfect occupation of O atoms, while imperfect occupation of Fe and P atoms leads to new (200) plane of FeP_4 phase which has a monoclinic crystal structure.

Keywords: LiFePO_4 , Imperfection, Synthesis, Solid-state, Calcination, XRD

1. Introduction

LiFePO_4 is a material that is widely used as a cathode active material for lithium-ion batteries [1-10]. To be able to be used as a cathode active material in lithium batteries, LiFePO_4 should have phospho-olivine structure [11]. An illustration of this structure can be seen in Figure 1. Basically, the phospho-olivine structure of LiFePO_4 is constructed by the bond of Fe, P and O. The arrangement of O atoms in phospho-olivine is almost close to a close-packed hexagonal structure [11]. Each of these O atoms is then bound to Fe atoms to form bonds that resemble FeO_6 , and some are bound to P atoms to form bonds that resemble PO-shaped tetrahedral [11]. Meanwhile, Li atoms are not bound to Fe, P, or O, so they are free to move and become Li^+ ions [11]. With these characteristics, the crystal structure of LiFePO_4 is composed of orthorhombic structural

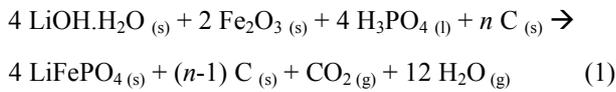
“frameworks” arranged by the arrangement of Li atoms, whereas in the orthorhombic structure there is an arrangement of phospho-olivine structures built from Fe-P-O (FeO_6 and PO_4) structures as illustrated in Figure 1 [11].

However, we often find the imperfect LiFePO_4 structures, especially in its phospho-olivine structure. This will probably result in poor electrochemical performance of lithium batteries when used as a cathode active material. LiFePO_4 itself can be synthesized through some conventional methods such as hydrothermal [12-16], sol-gel, [17] as well as solid-state reaction [18]. Of the various reaction methods, imperfections of the LiFePO_4 crystal structure are most often found in solid-state reaction path. This of course sparks its own interest to know the cause. It is however rare to find research that specifically addresses it. Therefore, this study was conducted to try to find out the common causes of imperfections in the LiFePO_4 crystal structures resulting from these solid-state reactions. This study is focused on the

crystal structure imperfections examined through x-ray diffraction (XRD) overview.

2. Methodology

In this study, a sample of LiFePO₄ powder synthesized by solid-state reaction was used as a case study sample. This sample was synthesized from raw materials i.e. LiOH.H₂O powder, Fe₂O₃ powder, and liquid H₃PO₄. These materials were mixed evenly then ground using mortar and pestle. After that, the sample was synthesized by the solid-state reaction method through three step calcinations i.e. 700°C for 2 hours, 800°C for 8 hours, and ended with inert calcination together with activated carbon tablets with the parameter of 800°C for 2 hours. Accumulatively, the whole reaction follows the following equation below:



The solid-state reaction then produces LiFePO₄ powder and residual of activated carbon tablets that do not react with oxygen. The LiFePO₄ powder was then separated from the remaining activated carbon tablets so that pure LiFePO₄ powder was obtained as a result of solid-state reaction. This LiFePO₄ powder was then used as a case study sample for this study.

This sample was then characterized by Rigaku SmartLab 3 kW X-ray Diffractometer (XRD) with Cu-tube ($\lambda = 1.541862 \text{ \AA}$) to investigate the imperfect crystal structure. The range of 2θ used is $10^\circ - 90^\circ$.

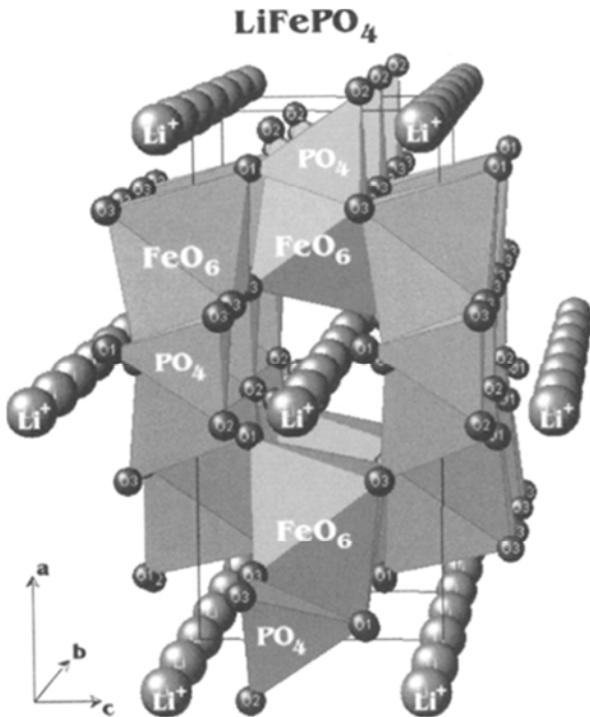


Figure 1. Illustration of phospho-olivine LiFePO₄ structure [11].

3. Results and Discussion

3.1. Precise Lattice Parameter Analysis

The list of 2θ and d (d -spacing) as important parameters of XRD characterization results can be seen in Table 1. This list showed matching d values when compared with the LiFePO₄ phase reference (PDF No. 01-080-6319), indicating the formation of LiFePO₄ with orthorhombic crystal structure (see Table 1).

However, to prove it, it is first necessary to do an analysis to identify the orthorhombic crystal structure in the sample of this case. An orthorhombic investigation was performed using Lutts' analytical method [19]. The result of Lutts' analytical method apparently did show the orthorhombic crystal structure ($a \neq b \neq c$, $\alpha = \beta = \gamma = 90^\circ$). The complete calculation result in the form of planes (hkl) and lattice parameters (a , b , c) can be seen in Table 2.

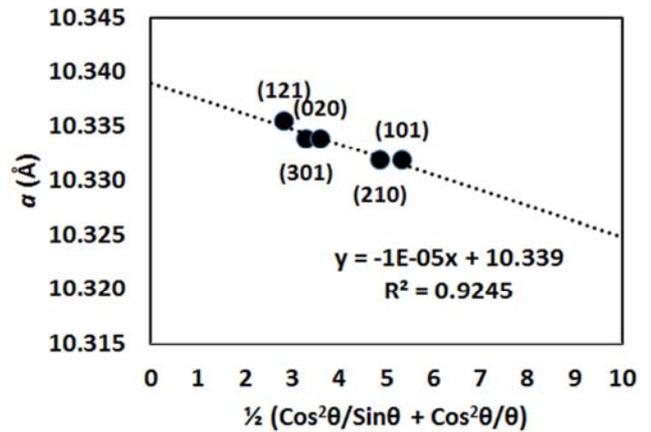


Figure 2. Linear regression analysis for calculating precise lattice parameter (a_0) of orthorhombic crystal structure.

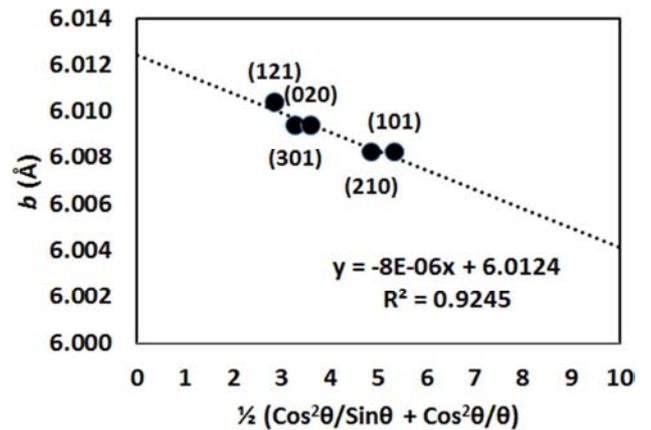


Figure 3. Linear regression analysis for calculating precise lattice parameter (b_0) of orthorhombic crystal structure.

Table 2 shows that orthorhombic lattice parameter values (a , b , and c) shows different values at different 2θ angle. This is because the lattice parameter values are pseudo-lattice parameter values. The value of this pseudo-lattice parameter varies depending on the 2θ point of view. Therefore, it is necessary to do an analysis to find out the true lattice

parameter values of the detected orthorhombic crystal structure. The true lattice parameter values can be calculated using the Nelson-Riley analytical technique [20]. The actual true lattice parameter value is then referred to as the precise lattice parameter value (symbolized by a_o , b_o , and c_o) [20]. Based on the distribution of pseudo-lattice parameter values in Table 2, the precise lattice parameter values are then calculated and the results are obtained in the form of linear regression equations as shown in Figure 2, Figure 3, and Figure 4.

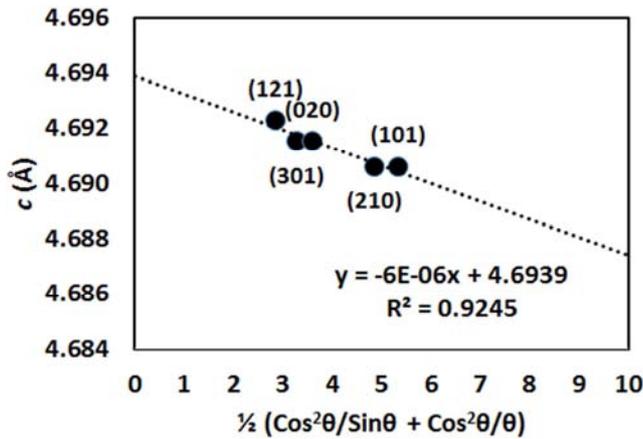


Figure 4. Linear regression analysis for calculating precise lattice parameter (c_o) of orthorhombic crystal structure.

In this precise lattice parameter analysis, the first thing to be noted is that the resulting linear regression line must be valid. The validity of linear regression line can be identified from the R^2 value. This value should ideally be 1 (one), but for real conditions it is sufficient to meet the range of $0.9 \leq R^2 \leq 1$ only. Equations plotted on Figure 2, Figure 3, and Figure 4 show the R^2 value of 0.9245, which means the resulting linear regression equation is valid for use. Furthermore, calculation results show that the orthorhombic crystal structure detected has precise lattice parameters of $a_o = 10.339 \text{ \AA}$, $b_o = 6.0124 \text{ \AA}$ and $c_o = 4.6939 \text{ \AA}$ (indicated by the intersection point on the y -axis in Figure 2, Figure 3, and Figure 4). These values possess close similarity with the lattice parameter values of LiFePO_4 phase reference (PDF No. 01-080-6319; $a = 10.32 \text{ \AA}$, $b = 6.0068 \text{ \AA}$, $c = 4.6993 \text{ \AA}$).

Next, we look back at Figure 2, Figure 3, and Figure 4. The linear regression lines formed are contributed by (101), (210), (020), (301), and (121) planes only. Other planes i.e. (111), (311), (222), and (512) do not contribute to the linear regression lines. A linear regression line basically shows an order of unit-cell of crystal structure, in this case orthorhombic crystal structure. Orthorhombic crystal structure formed in this case is only composed by (101), (210), (020), (301), and (121) planes only.

When traced deeper into the five constituent planes of this orthorhombic crystal structure of LiFePO_4 , these planes are significantly contributed by the arrangement of Li atoms, as seen on Figure 5. Thus, the orthorhombic crystal structure identified from the calculation of the precise lattice parameter is the orthorhombic “framework” structure composed by Li atoms.

The illustration of this orthorhombic structural “framework” can be seen in Figure 6.

Table 1. XRD characterization result and phase probability of LiFePO_4 .

XRD Characterization Result		Phase Reference of LiFePO_4 (PDF No. 01-080-6319)
2θ (°)	d (Å)	d (Å)
20.7980	4.271039	4.27677
22.7013	3.917075	3.91412
25.5950	3.480408	3.48394
29.7340	3.004687	3.00340
32.2407	2.776570	2.77577
35.6280	2.519976	2.51974
36.5570	2.458042	2.45787
52.4419	1.744843	1.74197
61.8502	1.500108	1.50170

Table 2. Lutts' analytical method result for orthorhombic analysis.

2θ (°)	hkl	a (Å)	b (Å)	c (Å)
20.7980	101	10.331766	6.008205	4.690591
22.7013	210	10.331768	6.008206	4.690592
25.5950	111	10.329693	6.006999	4.689650
29.7340	020	10.333776	6.009374	4.691504
32.2407	301	10.333803	6.009389	4.691516
35.6280	311	10.331519	6.008061	4.690479
36.5570	121	10.335455	6.010350	4.692266
52.4419	222	10.357234	6.023015	4.702154
61.8502	512	10.323952	6.003660	4.687044

3.2. Rietveld Refinement Analysis

The orthorhombic structure that has been calculated is known to be formed only by the arrangement of Li atoms only, while we know the structure of LiFePO_4 is of course also composed of Fe, P, and O atoms in addition to the Li atom. Since only orthorhombic “frame” structure of Li atoms has been identified, it can only be ascertained that there are Li atoms in (0, 0, 0) coordinates. Meanwhile, to ensure the formation of phospho-olivine structures within the “framework” of the orthorhombic structure, it is necessary to first ascertain the actual coordinates of the Fe, P, and O atoms. Furthermore, for knowing the actual coordinates of the Fe, P and O, we can do a simulation using the Rietveld refinement method [21].

This simulation is carried out until a convergent condition is reached between the observation results with the calculation of the Rietveld method. Besides achieving convergence, this simulation is also carried out until the lattice parameter values of a , b , and c are equal or close to the results of the a_o , b_o and c_o analysis results of the previous precise lattice parameter analysis. The result of the Rietveld refinement simulation can be seen in Figure 7. Convergence was achieved at $\chi^2 = 1.158$ and $wRp = 10.53\%$. Statistically, this condition is convergent enough to observe the simulation result. In this convergent condition, the lattice parameter values obtained are $a = 10.330258 \text{ \AA}$, $b = 6.005805 \text{ \AA}$ and $c = 4.695306 \text{ \AA}$. These are very close to the values of precise lattice parameters from previous calculation that is $a_o = 10.339 \text{ \AA}$, $b_o = 6.0124 \text{ \AA}$ and $c_o = 4.6939 \text{ \AA}$.

Table 3. Occupation coordinates of Li, Fe, P, and O atoms from Rietveld refinement simulation.

Atom	x/a	y/b	z/c	Occupation
Li	0.000000	0.000000	0.000000	1.0000
Fe	0.281317	0.250000	0.958766	0.8567
P	0.092807	0.250000	0.394131	0.4477
O	0.095833	0.250000	0.626202	1.0000
O	0.486515	0.250000	0.228455	1.0000
O	0.160234	0.050603	0.275679	1.0000

Table 4. Occupation coordinates of Li, Fe, P and O atoms of phospho-olivine LiFePO₄ reference (PDF No. 01-080-6319).

Atom	x/a	y/b	z/c	Occupation
Li	0.00000	0.00000	0.00000	1.0000
Fe	0.28222	0.25000	0.97472	1.0000
P	0.09486	0.25000	0.41820	1.0000
O	0.09678	0.25000	0.74279	1.0000
O	0.45710	0.25000	0.20602	1.0000
O	0.16558	0.04646	0.28478	1.0000

After convergent conditions are achieved, we can then pay attention to the simulation results of the coordinates/atom positions. These coordinates can be seen in Table 3. If we compare it with the atomic occupancy coordinates in the

reference of phospho-olivine LiFePO₄ structure (see table 4), we can see that only Li atoms have similar actual coordinates as the reference, unlike Fe, P, and O atoms. This further strengthens the notion that the ordered structure is formed by the arrangement of Li atoms only, which forms the orthorhombic “frame” structure. Meanwhile, the phospho-olivine structure which is expected to form on the inside of the orthorhombic structural space is thought to be still not visible because the coordinates of the Fe, P and O atoms are not in accordance with the standard reference of LiFePO₄ which has a phospho-olivine structure.

3.3. Predicted Causes of Phospho-olivine Structural Failure

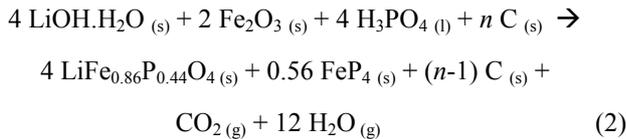
If we look at the occupancy values in Table 3, only Li and O atoms have full occupancy value of 1 (one); they are numbered in a unit-cell according to the LiFePO₄ standard reference (PDF No. 01-080-6319). Meanwhile, Fe and P atoms have occupancy values below 1 (one). This shows that the number of Fe and P atoms in an expected LiFePO₄ unit-cell is not in accordance with the reference of LiFePO₄ which has a phospho-olivine structure (PDF No. 01-080-6319).

Table 5. Phase probability of FeP₄ based on reference.

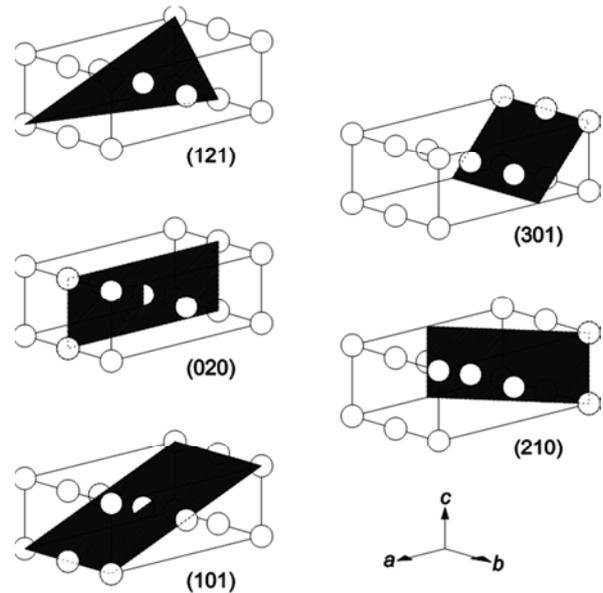
XRD Characterization Result		Phase Reference of FeP ₄ (PDF2 No. 79-0486)
2θ (°)	d (Å)	d (Å) and (hkl)
20.7980	4.271039	-
22.7013	3.917075	-
25.5950	3.480408	3.4820 (112)
29.7340	3.004687	3.0094 (023)
32.2407	2.776570	2.7614 (131)
35.6280	2.519976	2.5266 (200)
36.5570	2.458042	2.4428 (024)
52.4419	1.744843	1.7409 (224)
61.8502	1.500108	1.5016 ($\bar{1}$ 17)

If all Li, Fe, P and O atoms had an occupancy value of 1 (one), then the formula formed would be Li₁Fe₁P₁(O₄)₁ or LiFePO₄. However, in actual conditions, the occupancy values of Fe and P were 0.8567 and 0.4477 (see Table 3), resulting in the formula of LiFe_{0.8567}P_{0.4477}O₄. This shows that the LiFePO₄ formula failed to form due to the loss of Fe_(1-0.8567)P_(1-0.4477) or simplified to Fe_{0.1433}P_{0.5523}.

Fe_{0.1433}P_{0.5523} formula is very close to the Fe_{0.14}P_{0.56} formula, respectively. Fe_{0.14}P_{0.56} formula is actually 0.14 × FeP₄, so the prediction of the actual whole reaction that occurs is as follows:



From the predicted actual reaction (2) it is clearly seen that failure of phospho-olivine structure formation in LiFePO₄ is caused by incomplete reaction of unreacted Fe and P during solid-state synthesis. Instead of forming expected LiFePO₄, Fe and P bond together to form FeP₄. From phase point of view, FeP₄ has monoclinic crystal structure (PDF2 No. 79-0486).

**Figure 5.** Illustration of some planes constructed by Li atoms arrangement on orthorhombic “frame” of LiFePO₄ crystal structure.

Looking at XRD observation results, there are seven FeP_4 diffraction line probabilities (see Table 5). However, to ascertain the exact location of the FeP_4 crystal structure, it can be investigated from some anomalies in the Rietveld refinement error distribution. These anomalies are marked with number 1 and 2 in Figure 7.

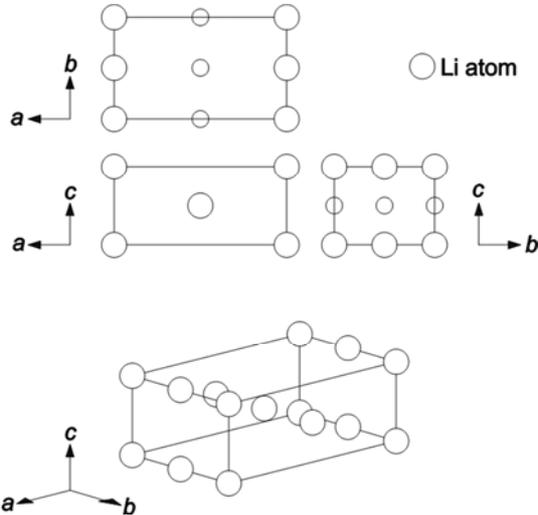


Figure 6. Illustration of Orthorhombic “frame” composed by arrangement of Li atoms.

These two anomalies are at $d = 2.519976 \text{ \AA}$ and $d = 5.16 \text{ \AA}$. At $d = 5.16 \text{ \AA}$; no convergence was found between observations and calculations result. This is caused by the non-detected value of $d = 5.16 \text{ \AA}$ on the observation results. The value of $d = 5.16 \text{ \AA}$ itself indicates the (200) plane in the phospho-olivine LiFePO_4 structure (PDF No. 01-080-6319). Thus, this indicates a failure in the formation of (200) plane in the resulting LiFePO_4 structure. The failure to form (200) plane is only possible due to imperfect occupancy of the O atoms. Meanwhile, the anomaly that occurs at $d = 2.519976 \text{ \AA}$ is probably caused by the addition of an additional preferred orientation from another phase beside LiFePO_4 . If we look at the probabilities in Table 5, in the range of d values there is a possible (200) plane of the FeP_4 phase which has a monoclinic crystal structure. From these two anomalous analysis, it can be assumed that the failure of the formation of phospho-olivine structure in the “framework” of orthorhombic structure is caused by the imperfections of occupational coordinates of Fe, P and O atoms, where imperfect coordinates of occupancy of the O atoms cause loss of (200) plane, which should be one of the planes forming the structure of phospho-olivine. Failure of occupancy of some Fe and P atoms consequently leads to the emergence of new (200) plane of another structure i.e. FeP_4 phase which has a monoclinic crystal structure.

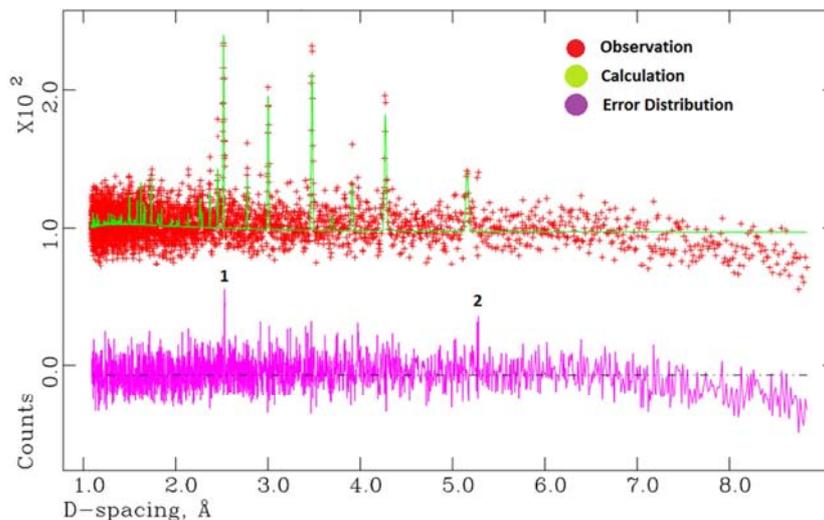


Figure 7. Rietveld refinement result of LiFePO_4 powder diffraction pattern.

4. Conclusion

Based on this XRD investigation, it can be concluded that there are two main factors causing imperfections of the structure of the phospho-olivine LiFePO_4 synthesized by solid-state reaction. The first factor is the imperfection in the occupancy coordinates of the O atoms, causing loss of (200) plane which should be one of the parts of the phospho-olivine structure. Whereas the second factor is occupancy failure of some Fe and P atoms which instead form a new crystal structure of monoclinic FeP_4 with a single (200) plane.

From the investigation of the LiFePO_4 powder as a case

study sample, although the structure of the phospho-olivine failed to form completely, the orthorhombic “frame” structure was formed perfectly by the arrangement of Li atoms.

Acknowledgements

This research is a part of the lithium battery research and development project at the Research Centre for Physics - LIPI funded by the Ministry of Research, Technology and Higher Education (Kemenristekdikti), Republic of Indonesia through the 2016 “Insinas” program.

References

- [1] N. D. Rago, J. Li, Y. Sheng, D. L. Wood, L. A. Steele, J. Lamb, C. Grosso, K. Fenton, I. Bloom. "Effect of Binder on the Overcharge Response in LiFePO₄-containing Cells," *Journal of Power Sources*, vol. 450, pp. 227595, Feb. 2020. DOI: 10.1016/j.jpowsour.2019.227595.
- [2] Q. Liang, H. Yue, S. Wang, S. Yang, K. H. Lam, X. Hou. "Recycling and Crystal Regeneration of Commercial Used LiFePO₄ Cathode Materials," *Electrochimica Acta*, vol. 330, pp. 135323, Jan. 2020. DOI: 10.1016/j.electacta.2019.135323.
- [3] X. Si, M. Li, X. Fu. "An Improved Method to Prepare FePO₄ by Introduction of Na₃PO₄ and its Usage for Fabricating LiFePO₄," *Journal of Wuhan University of Technology - Mater. Sci. Ed.*, vol. 34 (5), pp. 1097-1102, Oct. 2019. DOI: 10.1007/s11595-019-2165-0.
- [4] A. Kulka, K. Redel, J. Molenda. "Platelet-shape LiFePO₄/Fe₂P/C Composite Material as a High-rate Positive Electrode for Li-ion Batteries," *Solid State Ionics*, vol. 335, pp. 113-120, Jul. 2019. DOI: 10.1016/j.ssi.2019.03.004.
- [5] J. Li, J. Wu, Y. Li, H. Zhao, T. Zhao, S. Ma, H. Liu. "Facile Strategies to Utilize FeSO₄·7H₂O Waste Slag for LiFePO₄/C Cathode with High Performances," vol. 99, pp. 74-81, Jun. 2019. DOI: 10.1016/j.jtic.2019.03.002.
- [6] Y. Meng, Y. Li, J. Xia, Q. Hu, X. Ke, G. Ren, F. Zhu. "F-doped LiFePO₄@N/B/F-doped Carbon as High Performance Cathode Materials for Li-ion Batteries," *Applied Surface Science*, vol. 476, pp. 761-768, May 2019. DOI: 10.1016/j.apsusc.2019.01.139.
- [7] L. He, W. Zha, D. Chen. "Fabrication and Electrochemical Properties of 3D Nano-network LiFePO₄@Multiwalled Carbon Nanotube Composite Using Risedronic Acid as the Phosphorus Source," *Progress in Natural Science: Materials International*, vol. 29 (2), pp. 156-162, Apr. 2019. DOI: 10.1016/j.pnsc.2019.02.005.
- [8] S. X. Wu, C. L. Chiang, C. C. Wang, C. Y. Chen. "Functionalization of MWCNTs by Plasma Treatment and Use as Conductive Additives for LiFePO₄ Electrode," *Journal of the Taiwan Institute of Chemical Engineers*, vol. 89, pp. 208-214, Aug. 2018. DOI: 10.1016/j.jtice.2018.04.026.
- [9] G. Han, S. Yang, J. Liu, Y. Huang. "The Investigation of Humics as a Binder for LiFePO₄ Cathode in Lithium Ion Battery," *TMS 2018: Characterization of Minerals, Metals, and Materials 2018*, pp. 31-39, Jan. 2018. DOI: 10.1007/978-3-319-72484-3_4.
- [10] M. Talebi-Esfandarani, S. Rousselot, M. Gauthier, P. Sauriol, M. Duttine, A. Wattiaux, Y. Liu, A. X. Sun, G. Liang, M. Dollé. "Control of the LiFePO₄ Electrochemical Properties Using Low-cost Iron Precursor in a Melt Process," *Journal of Solid State Electrochemistry*, vol. 20 (12), pp. 3481-3490, Dec. 2016. DOI: 10.1007/s10008-016-3324-2.
- [11] M. Piana, M. Arrabito, S. Bodoardo, A. D'Epifanio, D. Satolli, F. Croce, B. Scrosati. "Characterization of Phospho-olivines as Materials for Li-ion Cell Cathodes," *Ionics*, vol. 8, pp. 17-26, 2002. DOI: 10.1007/BF02377749.
- [12] H. Yen, R. Rohan, C. Y. Chiou, C. J. Hsieh, S. Bolloju, C. C. Li, Y. F. Yang, C. W. Ong, J. T. Lee. "Hierarchy Concomitant in situ Stable Iron (II)-carbon Source Manipulation Using Ferrocenecarboxylic Acid for Hydrothermal Synthesis of LiFePO₄ as High-capacity Battery Cathode," *Electrochimica Acta*, vol. 253 (1), pp. 227-238, Nov. 2017. DOI: 10.1016/j.electacta.2017.09.065.
- [13] E. Golestani, M. Javanbakht, H. Ghafarian-Zahmatkesh, H. Beydaghi, M. Ghaemi. "Tartaric Acid Assisted Carbonization of LiFePO₄ Synthesized Through in Situ Hydrothermal Process in Aqueous Glycerol Solution," *Electrochimica Acta*, vol. 259 (1), pp. 903-915, Jan. 2018. DOI: 10.1016/j.electacta.2017.10.123.
- [14] R. Martins, R. Gonçalves, C. M. Costa, S. Ferdov, S. Lanceros-Méndez. "Mild Hydrothermal Synthesis and Crystal Morphology Control of LiFePO₄ by Lithium Nitrate," *Nano-Structures and Nano-Objects*, vol. 11, pp. 82-87, Jul. 2017. DOI: 10.1016/j.nanoso.2017.07.001.
- [15] G. Wu, N. Liu, X. Gao, X. Tian, Y. Zhu, Y. Zhou, Q. Zhu. "A Hydrothermally Synthesized LiFePO₄/C Composite with Superior Low-temperature Performance and Cycle Life," *Applied Surface Science*, vol. 435, pp. 1329-1336, Mar. 2018. DOI: 10.1016/j.apsusc.2017.11.276.
- [16] W. Zhou, C. Liu, Z. Wen, J. Xu, T. Han, G. Li, D. Huang, X. Liang, Z. Lan, H. Ning, H. Huang, J. Guo. "Effects of Defect Chemistry and Kinetic Behavior on Electrochemical Properties for Hydrothermal Synthesis of LiFePO₄/C Cathode Materials," *Materials Chemistry and Physics*, vol. 227, pp. 56-63, Apr. 2019. DOI: 10.1016/j.matchemphys.2019.01.050.
- [17] Y. Zhang, P. Xin, Q. Yao. "Electrochemical Performance of LiFePO₄/C Synthesized by Sol-gel Method as Cathode for Aqueous Lithium Ion Batteries," *Journal of Alloys and Compounds*, vol. 741, pp. 404-408, Apr. 2018. DOI: 10.1016/j.jallcom.2018.01.083.
- [18] M. Singh, B. Singh, M. Willert-Porada. "Reaction Mechanism and Morphology of the LiFePO₄ Materials Synthesized by Chemical Solution Deposition and Solid-state Reaction," *Journal of Electroanalytical Chemistry*, vol. 790, pp. 11-19, Apr. 2017. DOI: 10.1016/j.jelechem.2017.02.043.
- [19] A. Lutts. "A New Method for Determining the Lattice Parameters from Powder Patterns of Orthorhombic Crystals," *Zeitschrift für Kristallographie*, vol. 161 (3-4), pp. 195-207, Apr. 1982. DOI: 10.1524/zkri.1982.161.3-4.195.
- [20] J. B. Nelson, D. P. Riley. "An Experimental Investigation of Extrapolation Methods in the Derivation of Accurate Unit-cell Dimensions of Crystals," *Proc. Phys. Soc.*, vol. 57, pp. 160-177, 1945. DOI: 10.1088/0959-5309/57/3/302.
- [21] H. M. Rietveld. "A Profile Refinement Method for Nuclear and Magnetic Structures," *Journal of Applied Crystallography*, vol. 2, pp. 65-71, 1969. DOI: 10.1107/S0021889869006558.