

Synthesis, Characterization and Antibacterial Activity Evaluation of 4-amino Antipyrine Derivatives and Their Transition Metal Complexes

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Abstract: Due to their importance as catalysts in many reactions and their biological activities, an interest in the synthesis and characterization of transition metal complexes containing Schiff bases is increasing. Schiff base ligands have achieved considerable attention by the scientist over the decades as potential drug agent, Azomethine linkage (-CH=N-) of Schiff base play a significant role in medical chemistry. Derivatives of Schiff bases of 4-aminoantipyrine viz, 4-(2-hydroxy-3-methoxy benzylidene amino)-1,2-dihydro-2,3-dimethyl-1-phenylpyrazol-5-one (4) and 4-((5-methylfuran-2-yl) methylene amino)-1,2-dihydro-2,3-dimethyl-1-phenylpyrazol-5-one (5) and their Co (II), Ni (II), Cu (II) complexes were successfully synthesized, the Schiff bases ligand (4) and (5) were synthesized by condensation reaction. The structures of all the synthesized ligands were confirmed by using IR, UV-Visible, ¹H NMR, and ¹³C NMR. The Cu (II), Ni (II) and Co (II) complexes were confirmed by using IR and UV-Visible. The complexes are electrolytic in nature as indicated by molar conductance measurements. The data have shown that all complexes possess octahedral geometry. In-vitro antibacterial activity of all the synthesized ligands and their metal complexes were carried out by using disc diffusion method against *Staphylococcus aureus*, *Escherichia coli* and *Pseudomonas aeruginosa* bacterial strain. Compound Co (4) has exhibited better antibacterial activity than the standard drug against *S. aureus* (25 mm zone of inhibition compared to the standard antibiotic Oxacillin (23 mm zone of inhibition).

Keywords: Schiff Base, 4-aminoantipyrine, Ortho-vanillin, 5-methyl Furfural, Transition Metal Complexes, Antimicrobial

1. Introduction

Heterocyclic compounds are important class of organic compounds due to their chemical, biological, and technical significance, [1] these compounds are widely distributed in nature and essential to many biochemical processes [2].

Antipyrine (1,5-dimethyl-2-phenylpyrazole-3-one) is a compound that possesses a pyrazolone moiety with a five-membered lactam ring containing two nitrogens and a ketone in the same molecule [3, 4]. Antipyrine and its 4-amino derivative (4-amino-1,5-dimethyl-2-phenylpyrazole-3-one) forms a variety of Schiff bases with aldehydes, and a remarkable number of compounds have recently been reported [5, 6]. 4-Aminoantipyrine and its derivatives have been synthesized and investigated for its

biological activities such as anti-inflammatory [7, 8], antibacterial, [9, 10] anthelmintic [11], anticancer [12, 13] etc. activities

Schiff base compounds containing an azomethine group (-CH=N-) have received considerable attention in the field of chemistry and biology due to their chemotherapeutic value. [14, 15] Biological activity evaluation of transition metal complexes derived from Schiff base ligands were widely investigated and have attracted more attention by various researchers, schiff base ligands can coordinate to metal ions via azomethine nitrogen and have been studied extensively. In azomethine derivatives, the C=N linkage is essential for biological activity, several azomethine have been reported to possess remarkable antibacterial, antifungal, anticancer and antimalarial activities [16]

In this paper, we present the synthesis and anti bacterial activity investigations of Schiff bases of antipyrine and their metal complex against *Escherichia. coli*, *Pseudomonas aeruginosa* and *Staphylococcus aureus* bacterial strain. The major aim of this study was to compare the activity of the ligands with their metal complexes.

2. Result and Discussion

2.1. Antibacterial Activities

The antibacterial activity was determined with the disc diffusion method. The *in vitro* antibacterial screening was done against the bacterial species *Escherichia. coli*, *Pseudomonas aeruginosa* and *Staphylococcus aureus*. The bacterial cultures were inoculated in nutrient broth (inoculation medium) and incubated overnight at 37°C. Inoculated medium containing 24h grown culture was added aseptically to the nutrient medium and mixed thoroughly to get a uniform distribution. This solution was poured into Petri dishes and then allowed to attain room temperature. Wells (6 mm in diameter) were punched carefully using a sterile cork borer and were filled with test solution 10 mg/mL. The plate was allowed to stand for an hour in order to facilitate the diffusion of the drug solutions, then the plates were incubated at 37°C for 24h for bacteria and the diameter of the zone of inhibition were measured.

The synthesized Schiff bases and their Co (II), Ni (II), and Cu (II) complexes were tested in-vitro antibacterial activity against *Escherichia coli*, *Staphylococcus aureus*, and *Pseudomonas aeruginosa* by disc diffusion method. Oxacillin, tobramycin, and Cefotaxime were used as standard antibiotics drug. The zone of inhibition values indicates that all the compounds exhibited a varied range 6-25 mm of antibacterial activity (table 1) against all the tested bacterial strains.

Table 1. In-vitro antibacterial activity values of the synthesized ligands (4,5) and their Cu(II), Ni(II) and Co(II) complexes.

Antibacterial activity (inhibition zone diameter in mm)			
Compound	Bacterial strain		
	Staphylococcus aureus	Escherichia coli	Pseudomonas aeruginosa
4	6	6	6
Co (4)	25	22	14
Ni (4)	14	13	6
Cu (4)	19	12	6
5	6	6	6
Ni (5)	18	19	20
Cu (5)	15	14	6
Oxacillin	23		
Cefotaxime		32	
Tobramycin			21

Note: The size of 6 mm means that no activity (there was no zone of inhibition)

2.2. Molar Conductivity of the Metal Complexes

The molar conductance for the complexes measured in 10^{-3} M solution in dichloromethane as solvents at room temperature. The result, as shown in the table 2, demonstrated that the complexes have a binuclear nature.

Table 2. Molar conductance data of the complexes.

Compound	Λ ($\Omega^{-1}\text{cm}^2\text{mol}^{-1}$)	Nature
Cu complex of (4)	182	Electrolyte
Ni complex of (4)	68	Electrolyte
Co complex of (4)	56	Electrolyte
Cu complex of (5)	106	Electrolyte
Ni complex of (5)	62	Electrolyte

3. Experimental Section

3.1. Materials and Methods

All chemicals and solvents were purchased from Sigma-Aldrich and used without further purification; the reaction process was monitored by TLC silica gel plates; the purification of the products was performed using column chromatography using silica gel (100–200 mesh). The melting point was determined in a capillary tube using an electro thermal points apparatus and thermometer; Infrared spectra were recorded on Perkin-Elmer FT-IR model spectrophotometer in the range of 4000–400 cm^{-1} with a KBr pellet as a reference material; Electronic spectra were measured in Lambda 950 UV/VIS/NIR spectrometer in the range 200–800 nm at room temperature. ^1H and ^{13}C NMR were recorded using Bruker Avance NMR spectrometer (400 MHz) using $\text{DMSO}-d_6$ and CDCl_3 as solvent and tetramethyl silane (TMS) as an internal standard. Molar conductivities were recorded at room temperature using electrical conductivity meter. The antibacterial activity was determined with the Disc Diffusion method.

3.2. Chemistry

The titled compounds and their metal complexes described in this study were synthesized as outlined in figure 1, the Schiff base ligands were synthesized by the condensation of 4-aminoantipyrine with appropriate aldehyde in ethanol in the presence of acetic acid as a catalyst, the metal complexes of the ligands were prepared by refluxing the ligands with appropriate metal salt in ethanol.

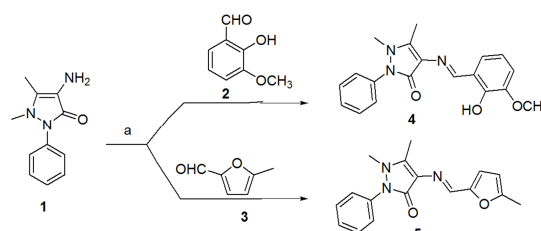


Figure 1. Synthesis of ligands (4) and (5): Reagent and condition: a) EtOH, AcOH, reflux.

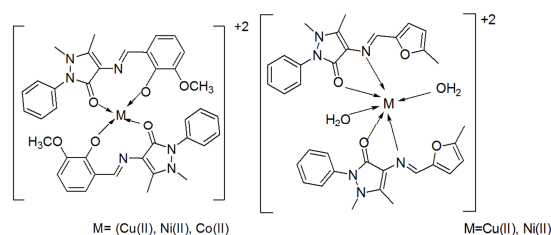


Figure 2. Proposed structure of the metal complexes.

3.3. Synthesis and Characterization of Ligands

3.3.1. General Procedure for the Synthesis of Schiff Bases (4) and (5)

A mixture of 4-aminoantipyrine (1) (0.5081g, 0.5 mmol) and appropriate aldehyde (2 or 3) (0.5 mmol) were dissolved in 20 mL of ethanol, 3-5 drops of glacial acetic acid was added as a catalyst. The reaction mixtures were refluxed at a temperature of 80°C for 6 hours. The progress of the reaction was monitored by TLC. After completion of the reaction, the resulting mixtures were cooled to room temperature and the obtained precipitates were filtered, washed several times with ethanol and dried over anhydrous calcium chloride under vacuum. Then the crude residue was further purified by column chromatography (Silica gel 200-400 mesh) using hexane: ethyl acetate as eluents affording the ligands (4 or 5)

3.3.2. Characterization of Schiff Bases of (4) and (5)

- i. 4-(2-hydroxy-3-methoxybenzylideneamino)-1,2-dihydro-2,3-dimethyl-1-phenylpyrazol-5-one (4)

Yellow crystal; Yield: 86%; Mp: 178-180°C; IR(KBr, Cm^{-1}): 3454(OH, phenolic), 1665(C=O, carbonyl), 1602(C=N, azomethine); $^1\text{H NMR}$ (CDCl_3 , 400MHz): δ 13.9(s, 1H, OH), 9.81(s, 1H, HC=N), 6.82-7.46(m, 8H, Ar-H), 3.86(s, 3H), 3.1(s, 3H), 2.37(s, 3H); $^{13}\text{C NMR}$ (CDCl_3 , 100MHz): δ 160.2, 150.3, 149.8, 148.0, 134.2, 129.3, 127.2, 124.6, 123.5, 120.1, 118.5, 115.8, 113.5, 110.1, 56.0, 35.5, 10.0

- ii. 4-((5-methylfuran-2-yl) methylene amino)-1,2-dihydro-2,3-dimethyl-1-phenylpyrazol-5-one (5)

Orange crystal; Yield: 80%; Mp: 179-180°C; IR(KBr, Cm^{-1}): 1638 (C=O, carbonyl), 1574(C=N, azomethine), 1492(HC=N); $^1\text{H NMR}$ (DMSO- d_6 , 400MHz): δ 9.80 (s, 1H, HC=N), 7.36-7.52 (m, 5H, Ar-H), 6.84 (d, J = 0.4Hz, furan proton), 6.26 (d, J = 0.4, furan proton), 3.43 (s, 3H), 2.38(s, 3H); $^{13}\text{C NMR}$ (DMSO- d_6 , 100MHz), δ 160.1, 155.2, 152.0, 151.9, 143.7, 134.9, 129.6, 127.3, 124.9, 117.0, 116.2, 109.2, 35.8, 14.3, 10.3

3.4. Synthesis and Characterization of Metal Complexes of (4)

3.4.1. Procedure for the Synthesis of Cu(II), Ni(II), and Co(II) Complexes of (4)

An ethanolic solution (10 ml) of (4) (0.337 gm, 0.5 mmol) was added drop wise to 5 ml hot methanol solution of Cu (CH_3COO) $_2$.H $_2$ O, Ni (CH_3COO) $_2$.H $_2$ O, or CoCl $_2$.6H $_2$ O (0.5mmol). The reaction mixture was refluxed at temperature of 70-80°C for 7-8 hours. The completion of the reaction was monitored by disappearance of spot on TLC plates using n-hexane: ethyl acetate (7:3) as mobile phase, the product obtained was filtered, washed thoroughly with ethanol. The sample was recrystallized using dichloromethane

3.4.2. Characterization of Metal Complexes of (4)

- i. Characterization of Copper (II) complexes of (4)

Fine Blue Black powder; Yield: 67%; Mp: 167-170°C; IR(KBr, Cm^{-1}): 1610(C=N, azomethine), 1657 (C=O, carbonyl). The imine bond C=N stretching shifted to higher value in the spectra of the complexes indicating the

coordination of azomethine nitrogen to the metal ions. The bands in the low frequency region 735 and 522 is due to the formation of $\nu(\text{M}-\text{O})$ and $\nu(\text{M}-\text{N})$ vibrations respectively.

Electronic spectrum of Cu complexes of (4)

The UV/visible electronic absorption spectra of the Schiff base ligand and its metal complexes were recorded in CDCl_3 at room temperature.

The spectrum of Cu (II) complex of ligand (4) shows three bands at 334, 358, and 430 nm. The two absorption band at 334 nm and 358 nm assigned to $\pi \rightarrow \pi^*$ and $n \rightarrow \pi^*$ transition respectively and band at 430 nm assigned to charge transfer

- ii. Characterization of Ni (II) complexes of (4)

Light green fine powder; Yield 72%; Mp: 140-142°C; IR(KBr, Cm^{-1}): 3418 (O-H) suggests that the coordinated water molecule is present in the Ni (II) complex, 1657 (C=O), 1583 (C=N, azomethine), 1456(CH=N), the bands in the low frequency region 726 and 561 is due to the formation of $\nu(\text{M}-\text{O})$ and $\nu(\text{M}-\text{N})$ vibrations respectively.

Electronic spectrum of Ni (II) complexes of (4)

The spectrum of Ni (II) complex shows three bands at 320, 334, and 366 nm. The bands at 320 and 334 nm assigned to carbonyl $\pi \rightarrow \pi^*$ and azomethine $n \rightarrow \pi^*$ transition respectively, and the bands at 366 nm could be assigned for corresponding charge transfer, which justified the octahedral geometry of Ni complex

- iii. Characterization of Co (II) complexes of (4)

Deep green Fine powder; Yield: 76.6%; Mp: 188-192°C; IR(KBr, Cm^{-1}): 1661(C=O, carbonyl), 1600 (C=N, azomethine), 1467(HC=N), the bands in the low frequency region 565 and 467 is due to the formation of (M-O) and (M-N) vibrations respectively

Electronic spectrum of Co(II) complexes of (4)

The electronic absorption spectral bands of Co (II) complex in chloroform are observed in the three main UV regions: 323, 335, and 370 nm. The electronic spectrum of Co (II) complex absorption bands at 323 nm is due to $\pi \rightarrow \pi^*$ transition of the carbonyl group (C=O). The band at 335 nm can be attributed to the $n \rightarrow \pi^*$ of the (HC=N) transitions. The peak at 370 nm is due to charge transfer.

3.5. Synthesis and Characterization of Metal Complexes of (5)

3.5.1. General Procedure for the Synthesis of Cu(II), Ni(II), Co(II) Complexes of (5)

An ethanolic solution (10 ml) of (5) (0.2953 gm, 0.5 mmol) was added drop wise to 5 ml solution of (0.5mmol) of the Cu (CH_3COO) $_2$.H $_2$ O, Ni (CH_3COO) $_2$.H $_2$ O in methanol with constant stirring. The reaction mixture was refluxed at temperature of 80°C, for 12 hours. The completion of the reaction was monitored by disappearance of spot on TLC plates using n-hexane: ethyl acetate (8:2) as mobile phase. The product obtained was filtered off, washed thoroughly with ethanol. The sample was recrystallized using dichloromethane

3.5.2. Characterization of Metal Complexes of (5)

- i. Characterization of Copper (II) complexes of (5)

Black sticky powder, Yield: 84.5%; Mp: 199-201°C; IR (KBr, cm^{-1}): 3418(OH, water molecules), 1657(C=O, carbonyl), 1602(C=N, azomethine), 1483(CH=N), the new bands in the low frequency region 735 and 552, is due to the formation of (M–O) and (M–N) vibrations respectively.

Electronic spectrum of Cu(II) complexes of (5)

The electronic absorption bands showed at 280, 296, and 343 nm, the bands at 296 nm is attributed to $n \rightarrow \pi^*$ transition of C=O groups and the band at 280 nm can be attributed to the

$\pi \rightarrow \pi^*$ transition of the HC=N functional group. The appearance of new bands at 343 nm is due to charge transfer

ii. Characterization of Ni (II) complexes of (5)

Light Brown Needle like Crystal, Yield: 70%; Mp: 135-138°C; IR(KBr, cm^{-1}): 3427(OH of water), 1657(C=O, carbonyl), 1565(C=N, azomethine), 1410(HC=N), new bands in the region of 680 and 452 is due to metal ligand (M–N) and (M–O) vibration respectively.

Electronic spectrum of Ni(II) complexes of (5)

The spectrum of Ni (II) Schiff base complex shows three bands at 248, 344, and 356, the band at 248 nm is due to $\pi \rightarrow \pi^*$ transition of the carbonyl group (C=O) and the band at 344 nm can be attributed to the $n \rightarrow \pi^*$ of the (HC=N) transitions, and the band at 356 nm assigned to charge transfer.

4. Conclusion

The synthesized Schiff base and their metal complex were tested in-vitro antibacterial activity against Escherichia coli, Staphylococcus aureus, and Pseudomonas aeruginosa by disc diffusion method using Oxacillin, tobramycin, and Cefotaxime as standard antibiotics drug compounds (4) and (5) exhibited no activity against E. coli, S. aureus, and P. aeruginosa. The complexes Cu(4) and Co(4) showed better activity against S. aureus and complexes of Ni(4), Ni(5), and Cu(5) showed moderate activity. The complexes Co (4) and Ni(5) showed better activity against E. coli but complexes Ni(4), Cu(4) and Cu(5) showed no activity against E. coli. The Complexes Ni (5) showed better activity against Pseudomonas aeruginosa and Co(4) showed moderate activity against Pseudomonas aeruginosa. But Ni(4) and Cu(4) do not showed any activity against Pseudomonas aeruginosa. In general the complexes show better activity than the ligands.

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